Periodic Table Practice



Directions: Please use the periodic table to fill in the missing information for each element.

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| **Be** | 76 | Radon | 65.38 | Silver |
| 23 | **Nd** | 98.907 | **Am** | Mercury |
| Antimony | 54.938 | **Mg** | 186.207 | 46 |
| 20.180 | Platinum | 100 | **Nb** | 17 |
| 2 | **Ta** | 83.80 | Americium | 54 |

Protons, Neutrons, and Electrons Practice

 How to calculate the number of each particle in an atom:

# Protons = Atomic Number

# Neutrons = Atomic Mass (rounded) – Atomic Number OR Big # - Small #

# Electrons = Protons

Use the periodic table to fill in the chart below.

|  |  |  |  |  |  |
| --- | --- | --- | --- | --- | --- |
| Element Symbol | Atomic Number | Atomic Mass | Protons | Neutrons | Electrons |
| H | 1 |  | 1 |  |  |
|  | 2 |  |  | 2 | 2 |
|  | 3 | 7 | 3 |  |  |
|  | 4 |  |  | 5 | 4 |
|  | 5 |  | 5 |  | 5 |
| C | 6 | 12 |  |  |  |
| N | 7 |  |  | 7 |  |
| O | 8 |  | 8 | 8 |  |
|  | 9 | 19 |  |  | 9 |
|  | 10 | 20 | 10 |  |  |
|  | 11 |  |  | 12 | 11 |
|  | 12 |  | 12 |  |  |
| Al | 13 |  |  |  | 13 |
|  | 14 |  | 14 | 14 |  |
| P | 15 |  |  |  | 15 |
| S | 16 |  | 16 |  |  |
|  | 17 | 35 |  |  | 17 |
|  | 18 |  | 18 | 22 |  |
|  | 19 | 39 |  |  | 19 |
| Ca | 20 |  | 20 |  |  |